

50 a. $C < Ca$ b. $Mg^{2+} < Fe^{2+}$ c. $F < Si$ d. $Zn^{2+} < Ca^{2+}$

56 a. I is smaller than Rb

Li is also smaller than Rb

This is tough to predict. The best way to answer this question is by looking at the values on the back of your periodic table.

b. Cl^- and Ca^{2+} are isoelectronic (They have identical electron configurations.) $Ca^{2+} < Cl^-$

58a. $Na < Ca < As$ b. $Sn < As < S$ c. $Ba < Sc < Zn < Al < Br < Kr$